PYQSide PYQSide.com — free PYQs with solutions Powered by PrepBlink

NEET STRUCTURE OF ATOM — PYQS

Max. Marks: 240 **Questions:** 60

Section A (MCQ) — 60 questions

Q.1 A bulb is rated at 150 watt, converting 8% energy into light. If energy of one photon is 4.42×10^{-19} J, how many photons are emitted by the bulb per second? [NEET (UG) 2026]

(1) 27.2×10^{19} (2) 4.06×10^{19}
(3) 1.35×10^{19} (4) 2.71×10^{19}

Scan or tap for the full solution

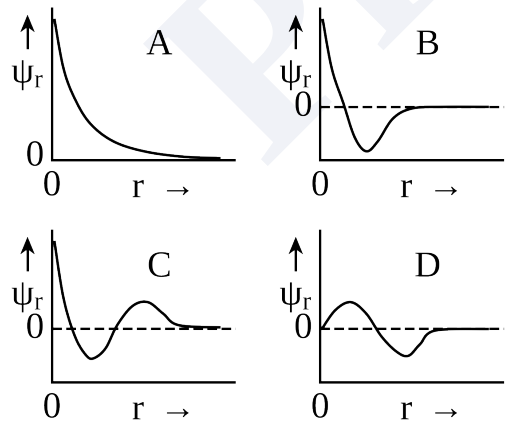
Q.2 Match **List-I** with **List-II** : [NEET (UG) 2026]

List-I (Quantum Numbers)		List-II (Orbital)	
A.	n = 2, l = 1	I.	3d
B.	n = 4, l = 0	II.	2p
C.	n = 5, l = 3	III.	4s
D.	n = 3, l = 2	IV.	5f

(1) A-II, B-III, C-IV, D-I
(2) A-I, B-II, C-III, D-IV
(3) A-II, B-III, C-I, D-IV
(4) A-IV, B-II, C-III, D-I

Solution

Q.3 Consider the following schematic plots of orbital wavefunction (ψ_r) against distance (r) from the nucleus.



The figure representing two radial nodes in the orbital is [RE-NEET (UG) 2026]

(1) D (2) A (3) B (4) C

Scan or tap for the full solution

Q.4 The ratio of the wavelengths of the light absorbed by a Hydrogen atom when it undergoes $n = 2 \longrightarrow n = 3$ and $n = 4 \longrightarrow n = 6$ transitions, respectively, is [NEET (UG) 2025]

(1) $\frac{1}{9}$ (2) $\frac{1}{4}$ (3) $\frac{1}{36}$ (4) $\frac{1}{16}$

Scan or tap for the full solution

Q.5 Energy and radius of first Bohr orbit of He^+ and Li^{2+} are [Given $R_H = 2.18 \times 10^{-18}$ J, $a_0 = 52.9$ pm] [NEET (UG) 2025]

(1) $E_n(Li^{2+}) = -19.62 \times 10^{-16}$ J;
$r_n(Li^{2+}) = 17.6$ pm
$E_n(He^+) = -8.72 \times 10^{-16}$ J;
$r_n(He^+) = 26.4$ pm
(2) $E_n(Li^{2+}) = -8.72 \times 10^{-16}$ J;
$r_n(Li^{2+}) = 17.6$ pm
$E_n(He^+) = -19.62 \times 10^{-16}$ J;
$r_n(He^+) = 17.6$ pm
(3) $E_n(Li^{2+}) = -19.62 \times 10^{-18}$ J;
$r_n(Li^{2+}) = 17.6$ pm
$E_n(He^+) = -8.72 \times 10^{-18}$ J;
$r_n(He^+) = 26.4$ pm
(4) $E_n(Li^{2+}) = -8.72 \times 10^{-18}$ J;
$r_n(Li^{2+}) = 26.4$ pm
$E_n(He^+) = -19.62 \times 10^{-18}$ J;
$r_n(He^+) = 17.6$ pm

Solution

Q.6 The energy of an electron in the ground state (n = 1) for He^+ ion is −x J, then that for an electron in (n = 2) state for Be^{3+} ion in J is [NEET (UG) 2024]

(1) −x (2) $-\frac{x}{9}$ (3) −4x (4) $-\frac{4}{9}x$

Scan or tap for the full solution

Q.7 Match **List I** with **List II**.

List I (Quantum Number)		List II (Information provided)	
A.	m_l	I.	Shape of orbital
B.	m_s	II.	Size of orbital
C.	l	III.	Orientation of orbital
D.	n	IV.	Orientation of spin of electron

Choose the correct answer from the options given below : [NEET (UG) 2024]

(1) A-I, B-III, C-II, D-IV

(2) A-III, B-IV, C-I, D-II

(3) A-III, B-IV, C-II, D-I

(4) A-II, B-I, C-IV, D-III

Solution

Q.8 The quantum numbers of four electrons are given below :

I. $n = 4; l = 2; m_l = -2; s = -\frac{1}{2}$

II. $n = 3; l = 2; m_l = 1; s = +\frac{1}{2}$

III. $n = 4; l = 1; m_l = 0; s = +\frac{1}{2}$

IV. $n = 3; l = 1; m_l = -1; s = +\frac{1}{2}$

The correct decreasing order of energy of these electrons is [RE-NEET (UG) 2024]

(1) IV > II > III > I (2) I > III > II > IV

(3) III > I > II > IV (4) I > II > III > IV

Scan or tap for the full solution

Q.9 Given below are two statements:

Statement-I: The Balmer spectral line for H atom with lowest energy is located at $\frac{5}{36}R_H$ cm^{-1}.

(R_H = Rydberg constant)

Statement-II: When the temperature of blackbody increases, the maxima of the curve (intensity and wavelength) shifts to shorter wavelength.

In the light of the above statements, choose the correct answer from the options given below: [RE-NEET (UG) 2024]

(1) Statement-I is true but Statement-II is false

(2) Statement-I is false but Statement-II is true

(3) Both Statement-I and Statement-II are true

(4) Both Statement-I and Statement-II are false

Scan or tap for the full solution

Q.10 Select the correct Statements from the following : Choose the correct answer from the options given below.

A. Atoms of all elements are composed of two fundamental particles.

B. The mass of the electron is 9.10939×10^{-31} kg.

C. All the isotopes of a given element show same chemical properties.

D. Protons and electrons are collectively known as nucleons.

E. Dalton's atomic theory, regarded the atom as an ultimate particle of matter. [NEET (UG) 2023]

(1) C, D and E only (2) A and E only

(3) B, C and E only (4) A, B and C only

Scan or tap for the full solution

Q.11 The relation between n_m (n_m = the number of permissible values of magnetic quantum number (m)) for a given value of azimuthal quantum number (l), is [NEET (UG) 2023]

(1) $l = 2n_m + 1$ (2) $n_m = 2l^2 + 1$

(3) $n_m = l + 2$ (4) $l = \frac{n_m - 1}{2}$

Scan or tap for the full solution

Q.12 Identify the incorrect statement from the following. [NEET (UG) 2022]

(1) All the five 5d orbitals are different in size when compared to the respective 4d orbitals.

(2) All the five 4d orbitals have shapes similar to the respective 3d orbitals.

(3) In an atom, all the five 3d orbitals are equal in energy in free state.

(4) The shapes of d_{xy}, d_{yz} and d_{zx} orbitals are similar to each other; and $d_{x^2-y^2}$ and d_{z^2} are similar to each other.

Scan or tap for the full solution

Q.13 If radius of second Bohr orbit of the He^+ ion is 105.8 pm, what is the radius of third Bohr orbit of Li^{2+} ion? [NEET (UG) 2022]

(1) 158.7 pm (2) 15.87 pm

(3) 1.587 pm (4) 158.7 Å

Scan or tap for the full solution

Q.14 Match **List-I** with **List-II** :

List-I (quantum number)		List-II (orbital)	
a.	n = 2, l = 1	i.	2s
b.	n = 3, l = 2	ii.	3s
c.	n = 3, l = 0	iii.	2p
d.	n = 2, l = 0	iv.	3d

Choose the correct answer from the options given below : [NEET (UG) 2022 (Phase 2)]

(1) (a)-(iii), (b)-(iv), (c)-(ii), (d)-(i)

(2) (a)-(iii), (b)-(iv), (c)-(i), (d)-(ii)

(3) (a)-(iv), (b)-(iii), (c)-(i), (d)-(ii)

(4) (a)-(iv), (b)-(iii), (c)-(ii), (d)-(i)

Solution

Q.15 When electromagnetic radiation of wavelength 300 nm falls on the surface of a metal, electrons are emitted with the kinetic energy of 1.68×10^5 J mol^{-1}. What is the minimum energy needed to remove an electron from the metal? ($h = 6.626 \times 10^{-34}$ J s, $c = 3 \times 10^8$ m s^{-1}, $N_A = 6.022 \times 10^{23}$ mol^{-1}) [NEET (UG) 2022 (Phase 2)]

(1) 2.31×10^5 J mol^{-1} (2) 2.31×10^6 J mol^{-1}
(3) 3.84×10^4 J mol^{-1} (4) 3.84×10^{-19} J mol^{-1}

Scan or tap for the full solution

Q.16 A particular station of All India Radio, New Delhi broadcasts on a frequency of 1,368 kHz (kilohertz). The wavelength of the electromagnetic radiation emitted by the transmitter is : [speed of light $c = 3.0 \times 10^8$ m s^{-1}] [NEET (UG) 2021]

(1) 21.92 cm (2) 219.3 m
(3) 219.2 m (4) 2192 m

Scan or tap for the full solution

Q.17 From the following pairs of ions which one is not an iso-electronic pair? [NEET (UG) 2021]

(1) Fe^{2+}, Mn^{2+} (2) O^{2-}, F^-
(3) Na^+, Mg^{2+} (4) Mn^{2+}, Fe^{3+}

Scan or tap for the full solution

Q.18 The number of protons, neutrons and electrons in $^{175}_{71}Lu$, respectively, are : [NEET (UG) 2020]

(1) 104, 71 and 71 (2) 71, 71 and 104
(3) 175, 104 and 71 (4) 71, 104 and 71

Scan or tap for the full solution

Q.19 How many electrons can fit in the orbital for which n = 3 and l = 1? [NEET 2016 (Phase 2)]

(1) 2 (2) 6 (3) 10 (4) 14

Scan or tap for the full solution

Q.20 4d, 5p, 5f and 6p orbitals are arranged in the order of decreasing energy. The correct option is : [NEET (UG) 2019]

(1) 6p > 5f > 4d > 5p (2) 5f > 6p > 4d > 5p
(3) 5f > 6p > 5p > 4d (4) 6p > 5f > 5p > 4d

Scan or tap for the full solution

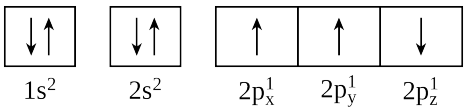
Q.21 Which of the following series of transitions in the spectrum of hydrogen atom falls in visible region? [NEET (UG) 2019]

(1) Paschen series (2) Brackett series
(3) Lyman series (4) Balmer series

Scan or tap for the full solution

Q.22 Which one is a wrong statement? [NEET 2018]

(1) Total orbital angular momentum of electron in s-orbital is equal to zero.
(2) An orbital is designated by three quantum numbers while an electron in an atom is designated by four quantum numbers
(3) The electronic configuration of N atom is



(4) The value of m for d_{z^2} is zero.

Scan or tap for the full solution

Q.23 Which one is the wrong statement ? [NEET 2017]

(1) The uncertainty principle is $\Delta E \times \Delta t \geq \frac{h}{4\pi}$
(2) Half filled and fully filled orbitals have greater stability due to greater exchange energy, greater symmetry and more balanced arrangement.
(3) The energy of 2s-orbital is less than the energy of 2p-orbital in case of hydrogen like atoms.
(4) de-Broglie's wavelength is given by $\lambda = \frac{h}{mv}$, where m = mass of the particle, v = group velocity of the particle.

Scan or tap for the full solution

Q.24 Two electrons occupying the same orbital are distinguished by [NEET 2016 (Phase 1)]

(1) azimuthal quantum number
(2) spin quantum number
(3) principal quantum number
(4) magnetic quantum number

Solution

Q.25 Which of the following pairs of d-orbitals will have electron density along the axes ? [NEET 2016 (Phase 2)]

(1) d_{z^2}, d_{xz} (2) d_{xz}, d_{yz}
(3) d_{z^2}, $d_{x^2-y^2}$ (4) d_{xy}, $d_{x^2-y^2}$

Scan or tap for the full solution

Q.26 Which is the correct order of increasing energy of the listed orbitals in the atom of titanium ? (At. no. Z = 22) [AIPMT 2015]

(1) 4s 3s 3p 3d (2) 3s 3p 3d 4s
(3) 3s 3p 4s 3d (4) 3s 4s 3p 3d

Scan or tap for the full solution

Q.27 The number of d-electrons in Fe^{2+} (Z = 26) is not equal to the number of electrons in which one of the following ? [AIPMT 2015 (Cancelled)]

(1) d-electrons in Fe (Z = 26)

(2) p-electrons in Ne (Z = 10)

(3) s-electrons in Mg (Z = 12)

(4) p-electrons in Cl (Z = 17)

Solution

Q.28 The angular momentum of electron in 'd' orbital is equal to [AIPMT 2015 (Cancelled)]

(1) $2\sqrt{3}\,\hbar$ (2) $0\,\hbar$

(3) $\sqrt{6}\,\hbar$ (4) $\sqrt{2}\,\hbar$

Scan or tap for the full solution

Q.29 What is the maximum number of orbitals that can be identified with the following quantum numbers?

$n = 3, l = 1, m_l = 0$ [AIPMT 2014]

(1) 1 (2) 2 (3) 3 (4) 4

Scan or tap for the full solution

Q.30 Calculate the energy in joule corresponding to light of wavelength 45 nm. (Planck's constant, $h = 6.63 \times 10^{-34}$ J s, speed of light, $c = 3 \times 10^{8}$ m s^{-1}) [AIPMT 2014]

(1) 6.67×10^{15} (2) 6.67×10^{11}

(3) 4.42×10^{-15} (4) 4.42×10^{-18}

Scan or tap for the full solution

Q.31 Be^{2+} is isoelectronic with which of the following ions? [AIPMT 2014]

(1) H^+ (2) Li^+ (3) Na^+ (4) Mg^{2+}

Scan or tap for the full solution

Q.32 What is the maximum number of electrons that can be associated with the following set of quantum numbers ?

$n = 3, l = 1$ and $m = -1$ [NEET 2013]

(1) 4 (2) 2 (3) 10 (4) 6

Scan or tap for the full solution

Q.33 Based on equation $E = -2.178 \times 10^{-18}\ J\left(\frac{Z^2}{n^2}\right)$, certain conclusions are written. Which of them is not correct ? [NEET 2013]

(1) Equation can be used to calculate the change in energy when the electron changes orbit.

(2) For n = 1, the electron has a more negative energy than it does for n = 6 which means that the electron is more loosely bound in the smallest allowed orbit.

(3) The negative sign in equation simply means that the energy of electron bound to the nucleus is lower than it would be if the electrons were at the infinite distance from the nucleus.

(4) Larger the value of n, the larger is the orbit radius.

Scan or tap for the full solution

Q.34 The value of Planck's constant is 6.63×10^{-34} J s. The speed of light is 3×10^{17} nm s^{-1}. Which value is closest to the wavelength in nanometer of a quantum of light with frequency of 6×10^{15} s^{-1} ? [NEET 2013]

(1) 50 (2) 75 (3) 10 (4) 25

Scan or tap for the full solution

Q.35 According to law of photochemical equivalence the energy absorbed (in ergs/mole) is given as ($h = 6.62 \times 10^{-27}$ ergs, $c = 3 \times 10^{10}$ cm s^{-1}, $N_A = 6.02 \times 10^{23}$ mol^{-1}) [NEET 2013 (Karnataka)]

(1) $\frac{1.196 \times 10^{8}}{\lambda}$ (2) $\frac{2.859 \times 10^{5}}{\lambda}$

(3) $\frac{2.859 \times 10^{16}}{\lambda}$ (4) $\frac{1.196 \times 10^{16}}{\lambda}$

Scan or tap for the full solution

Q.36 The orbital angular momentum of a p-electron is given as [AIPMT 2012 (Mains)]

(1) $\frac{h}{\sqrt{2}\pi}$ (2) $\sqrt{3}\frac{h}{2\pi}$

(3) $\sqrt{\frac{3}{2}}\frac{h}{\pi}$ (4) $\sqrt{6}\frac{h}{2\pi}$

Scan or tap for the full solution

Q.37 Maximum number of electrons in a subshell with l = 3 and n = 4 is [AIPMT 2012 (Prelims)]

(1) 14 (2) 16 (3) 10 (4) 12

Scan or tap for the full solution

Q.38 The correct set of four quantum numbers for the valence electron of rubidium atom (Z = 37) is [AIPMT 2012 (Prelims)]

(1) 5, 1, 1, $+\frac{1}{2}$ (2) 6, 0, 0, $+\frac{1}{2}$

(3) 5, 0, 0, $+\frac{1}{2}$ (4) 5, 1, 0, $+\frac{1}{2}$

Scan or tap for the full solution

Q.39 According to the Bohr theory, which of the following transitions in the hydrogen atom will give rise to the least energetic photon ? [AIPMT 2011 (Mains)]

(1) n = 6 to n = 1 (2) n = 5 to n = 4
(3) n = 6 to n = 5 (4) n = 5 to n = 3

Scan or tap for the full solution

Q.40 The total number of atomic orbitals in fourth energy level of an atom is [AIPMT 2011 (Prelims)]

(1) 8 (2) 16 (3) 32 (4) 4

Scan or tap for the full solution

Q.41 The energies E_1 and E_2 of two radiations are 25 eV and 50 eV respectively. The relation between their wavelengths i.e., λ_1 and λ_2 will be [AIPMT 2011 (Prelims)]

(1) $\lambda_1 = \lambda_2$ (2) $\lambda_1 = 2\lambda_2$
(3) $\lambda_1 = 4\lambda_2$ (4) $\lambda_1 = \frac{1}{2}\lambda_2$

Scan or tap for the full solution

Q.42 If n = 6, the correct sequence for filling of electrons will be [AIPMT 2011 (Prelims)]

(1) ns ⟶ (n–2)f ⟶ (n–1)d ⟶ np
(2) ns ⟶ (n–1)d ⟶ (n–2)f ⟶ np
(3) ns ⟶ (n–2)f ⟶ np ⟶ (n–1)d
(4) ns ⟶ np ⟶ (n–1)d ⟶ (n–2)f

Scan or tap for the full solution

Q.43 Maximum number of electrons in a subshell of an atom is determined by the following [AIPMT 2009]

(1) $2\ell + 1$ (2) $4\ell - 2$ (3) $2n^2$ (4) $4\ell + 2$

Scan or tap for the full solution

Q.44 A 0.66 kg ball is moving with a speed of 100 m/s. The associated wavelength will be
($h = 6.6 \times 10^{-34}$ J s) [AIPMT 2010 (Mains)]

(1) 6.6×10^{-32} m (2) 6.6×10^{-34} m
(3) 1.0×10^{-35} m (4) 1.0×10^{-32} m

Scan or tap for the full solution

Q.45 Which of the following is not permissible arrangement of electrons in an atom ? [AIPMT 2009]

(1) $n = 5,\ \ell = 3,\ m = 0,\ s = +1/2$
(2) $n = 3,\ \ell = 2,\ m = -3,\ s = -1/2$
(3) $n = 3,\ \ell = 2,\ m = -2,\ s = -1/2$
(4) $n = 4,\ \ell = 0,\ m = 0,\ s = -1/2$

Scan or tap for the full solution

Q.46 If uncertainty in position and momentum are equal, then uncertainty in velocity is [AIPMT 2008]

(1) $\frac{1}{m}\sqrt{\frac{h}{\pi}}$ (2) $\sqrt{\frac{h}{\pi}}$
(3) $\frac{1}{2m}\sqrt{\frac{h}{\pi}}$ (4) $\sqrt{\frac{h}{2\pi}}$

Scan or tap for the full solution

Q.47 The measurement of the electron position is associated with an uncertainty in momentum, which is equal to 1×10^{-18} g cm s^{-1}. The uncertainty in electron velocity is (mass of an electron is 9×10^{-28} g) [AIPMT 2008]

(1) 1×10^5 cm s^{-1} (2) 1×10^{11} cm s^{-1}
(3) 1×10^9 cm s^{-1} (4) 1×10^6 cm s^{-1}

Scan or tap for the full solution

Q.48 Consider the following sets of quantum numbers:

	n	l	m	s
(i)	3	0	0	+1/2
(ii)	2	2	1	+1/2
(iii)	4	3	−2	−1/2
(iv)	1	0	−1	−1/2
(v)	3	2	3	+1/2

Which of the following sets of quantum numbers are not possible? [AIPMT 2007]

(1) (i), (ii), (iii) and (iv) (2) (ii), (iv) and (v)
(3) (i) and (iii) (4) (ii), (iii) and (iv)

Scan or tap for the full solution

Q.49 The orientation of an atomic orbital is governed by [AIPMT 2006]

(1) principal quantum number
(2) azimuthal quantum number
(3) spin quantum number
(4) magnetic quantum number.

Solution

Q.50 Given : The mass of electron is 9.11×10^{-31} kg, Planck constant is 6.626×10^{-34} J s, the uncertainty involved in the measurement of velocity within a distance of 0.1 Å is [AIPMT 2006]

(1) 5.79×10^5 m s^{-1} (2) 5.79×10^6 m s^{-1}
(3) 5.79×10^7 m s^{-1} (4) 5.79×10^8 m s^{-1}

Scan or tap for the full solution

Q.51 The energy of second Bohr orbit of the hydrogen atom is −328 kJ mol^{-1}; hence the energy of fourth Bohr orbit would be [AIPMT 2005]

(1) −41 kJ mol^{-1} (2) −82 kJ mol^{-1}
(3) −164 kJ mol^{-1} (4) −1312 kJ mol^{-1}

Scan or tap for the full solution

Q.52 The frequency of radiation emitted when the electron falls from n = 4 to n = 1 in hydrogen atom will be (Given ionization energy of H = 2.18×10^{-18} J $atom^{-1}$ and h = 6.625×10^{-34} J s) [AIPMT 2004]

(1) 1.54×10^{15} s^{-1} (2) 1.03×10^{15} s^{-1}
(3) 3.08×10^{15} s^{-1} (4) 2.00×10^{15} s^{-1}

Scan or tap for the full solution

Q.53 The radioactive isotope $^{60}_{27}Co$ which is used in the treatment of cancer can be made by (n, p) reaction. For this reaction the target nucleus is [AIPMT 2004]

(1) $^{59}_{28}Ni$ (2) $^{59}_{27}Co$ (3) $^{60}_{28}Ni$ (4) $^{60}_{27}Co$

Scan or tap for the full solution

Q.54 The value of Planck's constant is 6.63×10^{-34} J s. The velocity of light is 3.0×10^{8} m s^{-1}. Which value is closest to the wavelength in nanometers of a quantum of light with frequency of 8×10^{15} s^{-1} ? [AIPMT 2003]

(1) 2×10^{-25} (2) 5×10^{-18}
(3) 4×10^{1} (4) 3×10^{7}

Scan or tap for the full solution

Q.55 Isoelectronic species are [AIPMT 2000]

(1) CO, CN^-, NO^+, C_2^{2-} (2) CO^-, CN, NO, C_2^-
(3) CO^+, CN^+, NO^-, C_2 (4) CO, CN, NO, C_2

Scan or tap for the full solution

Q.56 In hydrogen atom, energy of first excited state is −3.4 eV. Then find out K.E. of same orbit of hydrogen atom [AIPMT 2002]

(1) +3.4 eV (2) +6.8 eV
(3) −13.6 eV (4) +13.6 eV

Scan or tap for the full solution

Q.57 $^{235}_{92}U$, nucleus absorbs a neutron and disintegrates into $^{139}_{54}Xe$, $^{94}_{38}Sr$ and x so, what will be product x? [AIPMT 2002]

(1) 3-neutrons (2) 2-neutrons
(3) α-particle (4) β-particle

Scan or tap for the full solution

Q.58 The following quantum numbers are possible for how many orbitals :
n = 3, l = 2, m = +2 ? [AIPMT 2001]

(1) 1 (2) 2 (3) 3 (4) 4

Scan or tap for the full solution

Q.59 If a $^{a}_{b}X$ species emits first a positron, then two α and two β particles and in the last, one α is also emitted and gets converted to $^{c}_{d}Y$ species. So correct relation is [AIPMT 2001]

(1) c = a − 12, d = b − 5 (2) c = a − 5, d = b − 1
(3) c = a − 6, d = b − 0 (4) c = a − 4, d = b − 2

Scan or tap for the full solution

Q.60 For given energy, E = 3.03×10^{-19} Joules corresponding wavelength is (h = 6.626×10^{-34} J sec, c = 3×10^{8} m/sec) [AIPMT 2000]

(1) 65.6 nm (2) 6.56 nm
(3) 3.4 nm (4) 656 nm

Scan or tap for the full solution

ANSWER KEY

Ques.	1	2	3	4	5	6	7	8	9	10	11
Ans.	4	1	4	2	3	1	2	2	3	3	4
Ques.	12	13	14	15	16	17	18	19	20	21	22
Ans.	4	1	1	1	2	1	4	2	3	4	3
Ques.	23	24	25	26	27	28	29	30	31	32	33
Ans.	3	2	3	3	4	3	1	4	2	2	2
Ques.	34	35	36	37	38	39	40	41	42	43	44
Ans.	1	1	1	1	3	3	2	2	1	4	3
Ques.	45	46	47	48	49	50	51	52	53	54	55
Ans.	2	3	3	2	4	2	2	3	3	3	1
Ques.	56	57	58	59	60						
Ans.	1	1	1	1	4						