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NEET D- AND F-BLOCK ELEMENTS — POSITION AND ELECTRONIC CONFIGURATION OF D-BLOCK ELEMENTS — PYQS

Max. Marks: 12 **Questions:** 3

Section A (MCQ) — 3 questions

Q.1 Sc (Z = 21) is a transition element but Zn (Z = 30) is not because [NEET 2013 (Karnataka)]

(1) both Sc^{3+} and Zn^{2+} ions are colourless and form white compounds.

(2) in case of Sc, 3d orbitals are partially filled but in Zn these are filled.

(3) last electron is assumed to be added to 4s level in case of Zn.

(4) both Sc and Zn do not exhibit variable oxidation states.

Scan or tap for the full solution

Q.2 Which of the following ions has electronic configuration $[Ar]3d^6$?
(At. nos. Mn = 25, Fe = 26, Co = 27, Ni = 28)
[AIPMT 2010 (Prelims)]

(1) Ni^{3+} (2) Mn^{3+} (3) Fe^{3+} (4) Co^{3+}

Scan or tap for the full solution

Q.3 Among the following series of transition metal ions, the one where all metal ions have $3d^2$ electronic configuration is
[At. Nos. Ti = 22, V = 23, Cr = 24, Mn = 25]
[AIPMT 2004]

(1) Ti^{3+}, V^{2+}, Cr^{3+}, Mn^{4+} (2) Ti^{+}, V^{4+}, Cr^{6+}, Mn^{7+}

(3) Ti^{4+}, V^{3+}, Cr^{2+}, Mn^{3+} (4) Ti^{2+}, V^{3+}, Cr^{4+}, Mn^{5+}

Scan or tap for the full solution

ANSWER KEY

Ques.	1	2	3								
Ans.	**2**	**4**	**4**								