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NEET D- AND F-BLOCK ELEMENTS — MELTING POINT AND ENTHALPY OF ATOMISATION OF D-BLOCK ELEMENTS — PYQS

Max. Marks: 8 **Questions:** 2

Section A (MCQ) — 2 questions

Q.1 Given below are two statements: One is labelled as **Assertion A** and the other is labelled as **Reason R**.

In light of the above statements, choose the most appropriate answer from the options given below:
Assertion (A): Generally, 3d transition metals have high melting points.
Reason (R): Involvement of 3d-electrons in addition to 4s-electrons in the interatomic metallic bonding. [RE-NEET (UG) 2026]

(1) A is not correct but R is correct
(2) Both A and R are correct and R is the correct explanation of A
(3) Both A and R are correct and R is NOT the correct explanation of A
(4) A is correct but R is not correct.

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Q.2 Which one of the following does not correctly represent the correct order of the property indicated against it? [AIPMT 2012 (Mains)]

(1) Ti < V < Cr < Mn : increasing number of oxidation states
(2) $Ti^{3+} < V^{3+} < Cr^{3+} < Mn^{3+}$: increasing magnetic moment
(3) Ti < V < Cr < Mn : increasing melting points
(4) Ti < V < Mn < Cr : increasing 2nd ionization enthalpy

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ANSWER KEY

Ques.	1	2									
Ans.	2	3									