PART 1 - CHEMISTRY

Topic : Solid state

SECTION–1 : (Maximum Marks : 4)

- This section contains **ONE** question.
- The answer to each question is a **NUMERICAL VALUE**.
- For each question, enter the correct numerical value of the answer (in decimal notation, truncated / rounded-off to the **SECOND DECIMAL PLACE**; e.g. 6.25, 7.00, −0.33, −.30, 30.27, −127.30) in the place designated to enter the answer.
- Answer to each question will be evaluated according to the following marking scheme :

Full Marks : +4 If ONLY the correct answer is entered;
Zero Marks : 0 In all other cases.

1. The density (in g cm^{-3}) of the metal which forms a cubic close packed (ccp) lattice with an axial distance (edge length) equal to 400 pm is _____.
Use: Atomic mass of metal = 105.6 amu and Avogadro's constant = 6×10^{23} mol^{-1}

[JEE Advanced 2025 Paper 2]

Scan or tap for the full solution

Space for Rough Work

JEE ADVANCED SOLID STATE — DENSITY OF CUBIC UNIT CELL — PYQS

Test Pattern : **JEE (Advanced)**

ANSWER KEY

SECTION–1											
Ques.	1										
Ans.	**10.85 to 11.1**										