

PART 1 - CHEMISTRY

Topic : Electrochemistry

SECTION-1 : (Maximum Marks : 4)

- This section contains **ONE** question.
- The answer to each question is a **NON-NEGATIVE INTEGER**.
- For each question, enter the correct integer corresponding to the answer in the place designated to enter the answer.
- Answer to each question will be evaluated according to the following marking scheme :

Full Marks : +4 If ONLY the correct answer is entered;

Zero Marks : 0 In all other cases.

1. In an electrochemical cell, dichromate ions in aqueous acidic medium are reduced to Cr^{3+} . The current (in amperes) that flows through the cell for 48.25 minutes to produce 1 mole of Cr^{3+} is _____.

Use: 1 Faraday = 96500 C mol^{-1}

[JEE Advanced 2025 Paper 1]

Scan or tap for the
full solution

Space for Rough Work

Test Pattern : **JEE (Advanced)**

[illegible]