

PART 1 - CHEMISTRY

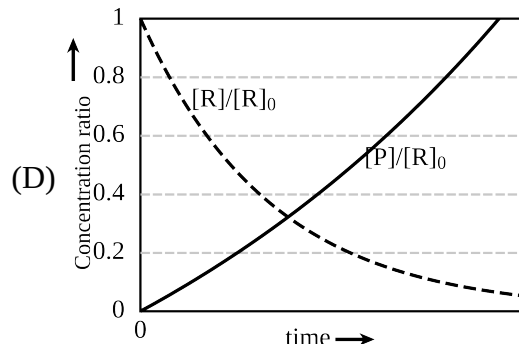
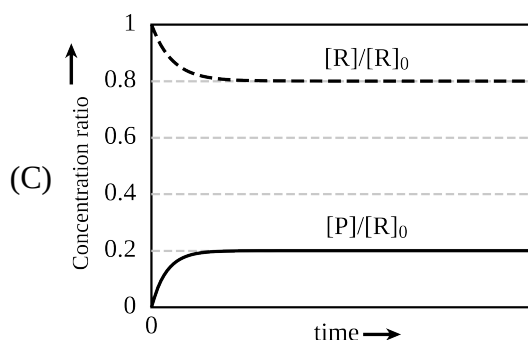
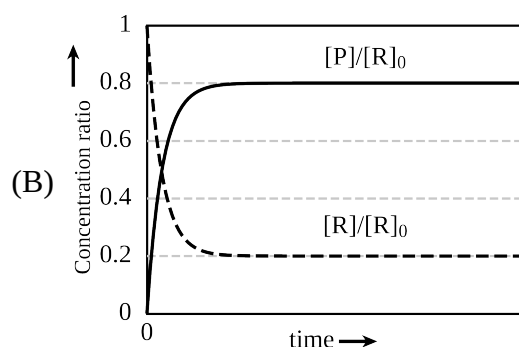
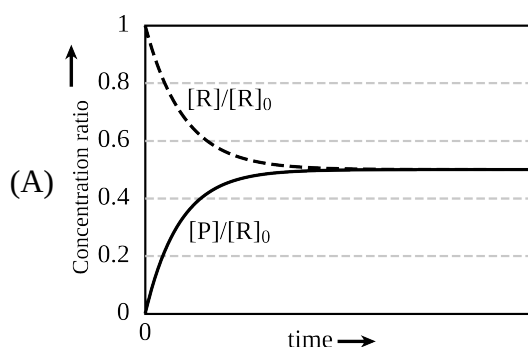
Topic : Chemical kinetics

SECTION-1 : (Maximum Marks : 3)

- This section contains **ONE** question.
- Each question has **FOUR** options (A), (B), (C) and (D). **ONLY ONE** of these four options is the correct answer.
- For each question, choose the option corresponding to the correct answer.
- Answer to each question will be evaluated according to the following marking scheme :
 Full Marks : +3 If **ONLY** the correct option is chosen;
 Zero Marks : 0 If none of the options is chosen (i.e. the question is unanswered);
 Negative Marks : -1 In all other cases.

1. For a reversible reaction $R \rightleftharpoons P$, at constant temperature, both the forward and the backward reactions are first order elementary reactions with rate constants k_f and k_b , respectively. At time zero, the concentration of R is $[R]_0$ and the concentration of P is zero. At any given time, $[R]$ and $[P]$ are the concentrations of R and P, respectively. If $k_b = 4k_f$, the correct graphical representation of the reaction is

[JEE Advanced 2026 Paper 1]



Scan or tap for the full solution



Space for Rough Work

Test Pattern : **JEE (Advanced)**

SECTION-1

[illegible]